

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## Home Work 401: Intro to Periodic Table

- \_\_\_\_\_ In his periodic table, Mendeleev did not always list elements in order of increasing atomic mass because he wanted to group elements with similar
  - Properties
  - Atomic numbers
  - Densities
  - Colors
- \_\_\_\_\_ Mendeleev arranged the elements by increasing atomic:
  - Number
  - Mass
  - Color
  - Density
- \_\_\_\_\_ A few elements were out of place on Mendeleev's periodic table. Name the scientist who put them in the right place
  - Mendeleev
  - Democritus
  - Moseley
  - Dalton
- \_\_\_\_\_ Mosely discovered that elements with similar properties occurred at regular intervals when arranged in order of increasing...
  - Atomic mass
  - Density
  - Color
  - Atomic number
- \_\_\_\_\_ What is it about the elements in a column of the periodic table that makes them behave similarly?
  - Atomic number
  - Atomic mass
  - Electron distribution
  - Radioactivity
- \_\_\_\_\_ The discovery of the noble gases changed Mendeleev's periodic table by adding a:
  - Period
  - Series
  - Group
  - Sublevel
- \_\_\_\_\_ Then modern periodic law states that:
  - No two electrons can have the same spin in the same atomic orbital
  - The physical and chemical properties of the elements are functions of their atomic number
  - Electrons exhibit properties of both particles and waves
  - The chemical properties of elements can be grouped according to periodicity, by physical properties cannot
- \_\_\_\_\_ Which of the following is the noble gas distribution for arsenic?
  - $[Kr] 4s^2 3d^{10} 4p^3$
  - $[Ar] 4s^2 3d^{10} 4p^3$
  - $[Ge] 4p^3$
  - $[K] 4s^1 3d^{10} 4p^3$